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Thermodynamic consideration of the retention mechanism in a poly(perfluoroalkyl ether) gas chromatographic stationary phase used in packed columns

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Abstract

Retention volumes of fifteen hydrocarbons were measured in columns containing several concentrations of a commercial poly(perfluoroalkyl ether), Fomblin Y HVAC 140/13, as the stationary phase. Two different types of packing were studied: one of them employed pre-silylated Chromosorb P AW DMCS as the solid support and the other type was prepared by coating the stationary phase on Chromosorb P AW and silylating on-column. On-column silylated columns showed unequivocal symptoms of partial deactivation; retention volumes changed regularly with the content of stationary phase in pre-silylated columns. Analysis of retention in pre-silylated columns indicates that a mixed mechanism (partition and adsorption onto the gas–liquid interface) is operative. The systems hydrocarbon–perfluorocompound show pronounced positive deviations from the ideal behaviour that can be attributed to repulsion between the hydrocarbon and the perfluorocompound segments.

1. Introduction

Perfluorinated substances have been sporadically employed as stationary phases in gas chromatography, and this mostly for the separation of substances of high chemical reactivity or perfluorocompounds and freons. Excellent review articles have been written by Pomaville and Poole [1] and by O'Mahony et al. [2].

Highly fluorinated fluids are characterized by their chemical inertia and by very weak molecular interactions, both in the pure state and in mixtures with other substances. Cohesive energy densities of perfluoroalkanes and perfluoro-

cycloalkanes are much lower than those of the corresponding hydrocarbons [3], and although fluorocarbons and hydrocarbons are individually highly non-polar, their mixtures show important deviations from Raoult's law [3–5]. These properties have two important consequences in relation with their use as gas chromatographic stationary phases. On the one hand, as pointed out by Poole and co-workers [1,6,7], retention times in highly fluorinated stationary phases will be shorter than in conventional phases, thus enabling the separation of low-volatility or thermally labile substances. On the other hand, weak interactions with the solid support or the capillary wall result in insufficient surface deactivation, uneven distribution and poor film stability,

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reflected in peak asymmetry, low efficiency and retention time variations; all of these deleterious symptoms are very markedly displayed when using perfluoroparaffins.

Drastic improvements in the chromatographic behaviour are obtained by incorporating polar groups in the stationary phase molecules. The poly(perfluoroalkyl ether) phase Fomblin YR,

$[(OCFCF_3CF_2)_n-(OCF_2)_m]$, was introduced in 1983 by Dhanesar and Poole [6,7]; paraffinic and olefinic hydrocarbons, injected at temperatures markedly lower than their boiling points, elute from these columns with short retention times, with very good peak shapes, but asymmetric peaks are obtained for solutes of higher polarity. Furthermore, the authors found that packings prepared by coating previously silylated supports (Gas-Chrom Q) gave symmetrical peaks when tested at low temperatures, but after heating to above 100°C and returning to low temperatures for re-testing, shorter retention times and asymmetric peaks were obtained, a behaviour that the authors attribute to film contraction, which results in bare-surface exposition and liquid lenses. Columns stable at temperatures higher than 200°C and with efficiencies similar to those obtained for conventional phases were obtained by coating Chromosorb P, packing the columns with coated support, and on-column silylating by injection of Silyl-8 or bis(trimethylsilyl)trifluoroacetamide (BSTFA) [6,7].

Mixtures of fluorocarbon compounds and aliphatic or alicyclic hydrocarbons are the only known examples of systems with large positive excess functions where none of the components is polar; as such, they have been the subject of many studies since the 1950s [8,9]. Although these trends are qualitatively consistent with the large differences in cohesive energy densities of fluorocarbons and hydrocarbons, the solubility parameter approach does not give a satisfactory quantitative explanation [3]. Plots of surface tension against composition also show pronounced deviations from the ideal behaviour [10,11]. The objective of the present paper is to obtain thermodynamic data for the processes of solution and adsorption on the gas-liquid interface for a group of hydrocarbons at infinite

dilution in a highly fluorinated stationary phase near room temperature. A high-molecular-mass perfluoroalkane was supposed to be used as the solvent, according with the original plan; the poor performance of these substances as stationary phases compelled us to modify the project. Fomblin was then chosen because of its characteristics and to profit from the experience collected by Poole and co-workers on this stationary phase.

2. Experimental

2.1. Materials and columns

Fomblin Y HVAC 140/13 (weight average molecular mass 6500) and 1,1,2-trichlorotrifluoroethane (Freon 113) were purchased from Aldrich; BSTFA was obtained from Supelco. Hydrocarbon solutes of different origins, all of them more than 99% pure, were used as received. Column packings were prepared in a rotary evaporator, using Freon 113 as solvent. Coated supports were packed into 1.0 m × 0.53 cm I.D. stainless-steel tubes with the aid of vacuum suction and gentle tapping. Two sets of columns were prepared:

(a) On-column silylated columns: a packing containing 5.67% by weight of Fomblin on Chromosorb P AW 60/80 was prepared by the above-mentioned method. A 1.80 m × 0.4 cm I.D. glass column containing this packing was installed in a Konik gas chromatograph (see later) with a glass-lined injector; the injector and the column oven were set to 100 and 150°C, respectively, and on-column silylation was performed by slowly injecting 100 μl of BSTFA with a nitrogen flow-rate of 5 ml/min. The column was held at 150°C for several hours, and the silylating treatment was then repeated after reversing the column end connections. The column was carefully emptied and its content was used as starting material to prepare packings with higher stationary-phase concentrations, by successively adding Fomblin by the above-mentioned coating technique. By this method four columns containing 5.67, 8.30,

13.36 and 22.05 weight percentage (*w*) of Fomblin on the same lot of solid support were prepared, with the objective of minimizing unpredictable effects associated with the silylating process.

(b) Pre-silylated columns, containing 4.29, 5.56, 7.50, 10.89, 15.52 and 19.40% by weight of Fomblin were prepared using Chromosorb P AW DMCS 60/80 as the solid support.

2.2. Apparatus and procedure

The study of thermal stability of packings was performed in a Konik 3000 gas chromatograph, equipped with a flame ionization detector (FID) and a Spectra Physics Datajet Integrator, using 1.50×2 mm I.D. glass columns.

A home-assembled apparatus, in which column temperature was controlled to better than $\pm 0.05^\circ\text{C}$ by immersion in a water bath, was used for the remaining measurements. Nitrogen, successively passed through a molecular sieves trap (Davidson 5A), a Brooks 8606 pressure regulator, a Brooks 8743 flow controller and a $2 \text{ m} \times 1/8$ in. O.D. coiled copper tube immersed in the column bath, was used as the carrier gas. Inlet pressures were measured by means of a mercury manometer at a point between the copper coil and a Swagelok $1/4$ in. s.s. "T"; one branch of the latter was provided with a septum, and the columns were connected to the remaining branch. Solute vapours were on-column injected by using Hamilton syringes. Eluates were detected with a Hewlett-Packard 5750 FID and electrometer signals were fed to a Hewlett-Packard 3396A integrator. Flow-rates ranging between 15 and 30 ml/min were measured by means of a water jacketed soap film flow meter.

Sample sizes of the order of $10^{-2} \mu\text{mol}$ resulted in symmetrical peaks for all the studied solutes, thus warranting Henry's law conditions for all the retention mechanisms. Aromatic hydrocarbons, whose peaks displayed a slight asymmetry that persisted at the smallest sample sizes compatible with instrumental noise, were excluded from this study. Solute vapours and a small methane sample were simultaneously in-

jected; net retention times were measured to 10^{-3} min between the maxima of the solute and the methane peaks. Specific retention volumes (V_g^0) and net retention volumes per gram of packing (V_N) were calculated from values of the operating parameters in the usual form [12]. Values at each temperature were means of no less than four injections; retentions were measured for groups of three to four solutes at five temperatures equally spaced within the range $22\text{--}35^\circ\text{C}$, and the measurements for the same group of solutes were repeated after two to three weeks.

Densities of Fomblin were measured at twelve temperatures between 19 and 41°C with a 3-ml pycnometer that had been carefully calibrated through the same temperature interval, and were least-squares fitted to the following polynomial:

$$\rho_2(T/^\circ\text{C}) = 1.94429 - 1.2808 \cdot 10^{-3}t - 1.30 \cdot 10^{-5}t^2 + 10^{-7}t^3 \quad (1)$$

The thermal expansion coefficient, as calculated from Eq. 1, is $\alpha_2 = 9.15 \cdot 10^{-4} \text{ K}^{-1}$ at 25°C .

3. Results and discussion

3.1. Comparing pre-silylated and on-column silylated packings

The pattern displayed in Fig. 1 for four representative cases is common for all the solutes. V_N increases in a satisfactory linear fashion with percentage Fomblin in the pre-silylated Chromosorb P phases. An important decrease in V_N occurs, particularly with unsaturated solutes, when the Fomblin proportion increases above 5.7%, on post-silylated packings.

A series of alternative suggestions could be made to explain the behaviour observed in on-column silylated packings. Identification of the retention mechanisms that operate in these columns and discriminating between their respective contributions to the total retention would be a difficult task, with dubious results. Pre-silylated packings represent a more attractive option

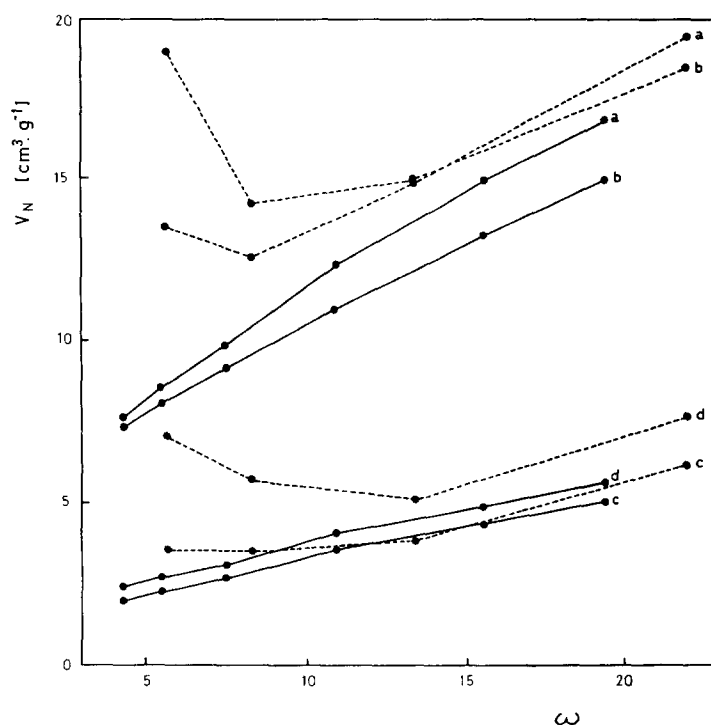


Fig. 1. Retention volume per gram of packing at 25°C (V_N) against percentage by weight of stationary phase (w). Dashed lines: on-column silylated packings. Solid lines: pre-silylated packings. Solutes: (a) *n*-octane; (b) 1-octene; (c) cyclohexane; (d) cyclohexene.

when the objective is to isolate the effects of the liquid phase from those of the solid support.

The physical meaning of retention volumes measured in pre-silylated columns, on the other hand, could be questioned in terms of the thermal instability of these packings. Therefore, and although the retention volumes demonstrated excellent reproducibility in the 22–35°C temperature range, a test addressed to check the stability of pre-silylated packings under conditions considerably more stringent than those prevailing during the thermodynamic measurements was performed. *n*-Nonane and *n*-decane retention volumes were measured at six temperatures between 40 and 75°C in columns containing 10.89 and 19.40% packings; the columns were then slowly heated to 120°C and kept for 1.5 h at this temperature, spontaneously cooled down to room temperature, and heated again to 120°C for 1.5 h; retention volumes were then re-measured in the 40–75°C interval. The results of this

experiment are summarized in Table 1, and show not only that the retention volumes measured before and after heating are coincident, but also that the slopes of the plots of $\ln V_g^0$ versus $1/T$ remain constant. The efficiency of both columns is poor and heating seems to produce a drop in the number of theoretical plates of the column containing the 10.89% packing. Summarizing, there are no counter-indications for the employment of these columns in the measurement of thermodynamic properties, where column efficiency is a desirable but not an essential parameter.

Only results obtained using pre-silylated packings are referred to in the following sections.

3.2. Obtaining the thermodynamic results

Experimental results obtained on each column were fitted to the equation

Table 1

Effect of thermal treatment on specific retention volumes, their temperature dependence and the efficiency of columns packed with Fomblin coated on Chromosorb P AW DMCS

	<i>n</i> -Decane		<i>n</i> -Nonane	
	Pre ^a	Post ^a	Pre ^a	Post ^a
<i>w</i> = 10.89%				
V_g^0 , 45°C	193.7	192.7	93.75	94.00
V_g^0 , 55°C	129.9	129.2	65.32	65.64
V_g^0 , 70°C	67.56	67.37	36.53	36.63
B^b	4626	4663	4149	4170
<i>N</i> , 50°C ^c	1670	1450	1450	1200
<i>N</i> , 70°C ^c	1400	990	875	810
<i>w</i> = 19.40%				
V_g^0 , 50°C	115.5	113.4		
V_g^0 , 62°C	70.82	70.00		
V_g^0 , 74°C	44.64	44.82		
B^b	4506	4409		
<i>N</i> , 50°C ^c	1260	1320		
<i>N</i> , 74°C ^c	1300	1260		

^a Results before and after the thermal treatment (details in the text).

^b Least-squares slopes for the fit to the equation $\ln V_g^0 = B/T + \text{constant}$.

^c Number of theoretical plates.

$$\ln V_g^0 = -\Delta H_s^0/RT + \text{constant} \quad (2)$$

where ΔH_s^0 , the heat of sorption, corresponds to the transfer of one mole of solute from an ideal vapour phase at a partial pressure of 1 atm to a stationary phase of yet undefined characteristics. Regression analysis of the results obtained in a given column showed that: (a) values of percent standard deviation of the heats of sorption, $100\sigma(\Delta H_s^0)/\Delta H_s^0$, were smaller than 1.5% for all the solutes; (b) retention volumes obtained in a just conditioned column differed by less than 0.8% from those obtained under identical conditions after two or more weeks of column use, and the differences between the corresponding heats of sorption were smaller than 4%.

In Figs. 2 and 3 V_g^0 values at 25°C and heats of sorption, respectively, have been plotted against *w* for a group of representative solutes. The existence of a mixed retention mechanism is made evident by both plots, since straight lines parallel to the horizontal axis should be obtained for a pure partitioning process; furthermore, it

can be deduced from Fig. 3 that the adsorptive component is markedly more exothermic than its solution counterpart [13]. Several years ago Conder et al. [14] proposed the following equation to express the retention volume per gram of packing when mixed mechanisms are operative:

$$V_N = K_L V_L + K_A A_L + K_S A_S \quad (3)$$

where K_L is the liquid–gas partition coefficient, K_A and K_S are the adsorption coefficients at the gas–liquid and at the liquid–solid interfaces, respectively, and V_L , A_L and A_S represent the liquid volume, the gas–liquid interfacial area and the liquid–solid interfacial area, all of them expressed per gram of packing; adsorption on bare portions of the support is not considered in Eq. 3, which seems reasonable for packings containing more than 3–4% by weight of stationary phase on Chromosorb P. Models endeavouring more detailed descriptions, involving equations with larger number of terms and, as such, of more difficult and dubious applicability to real data, have been more recently proposed [15].

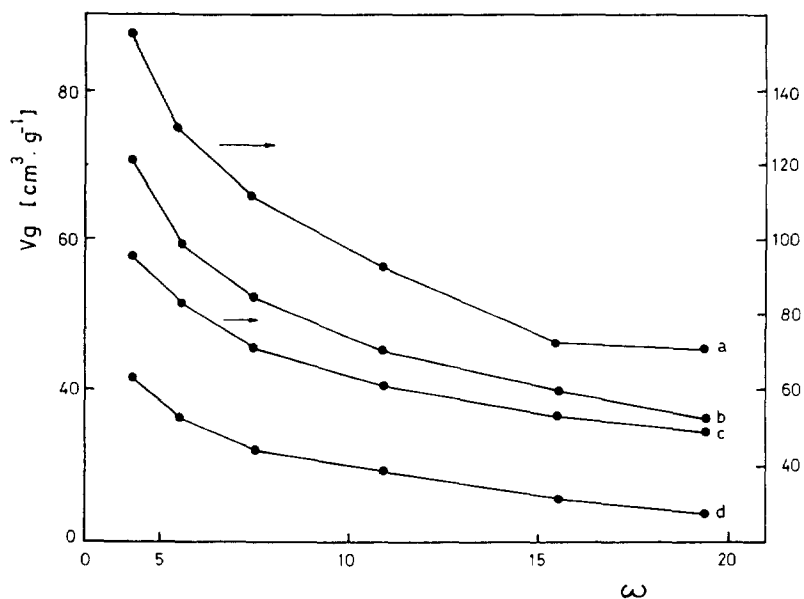


Fig. 2. Specific retention volumes at 25°C (V_g^0) against percentage by weight of stationary phase (w). Solutes: (a) 1-octene; (b) *n*-heptane; (c) 2,2-dimethylhexane; (d) cyclohexene.

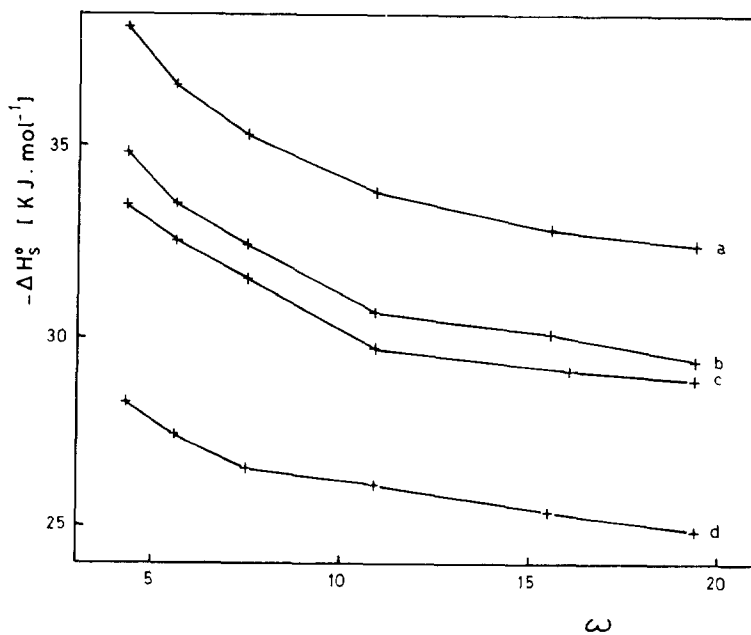


Fig. 3. Heats of sorption (ΔH_s^0) against percentage by weight of stationary phase (w). Solute symbols as in Fig. 2.

The values of A_L necessary to fit experimental results to Eq. 3 were calculated by following the proposal of Martire et al. [16], taking into consideration the large difference between the densities of β, β' -thiodipionitrile and Fomblin [17].

Multiple linear regression to Eq. 3, assigning to A_S values calculated from the specific surface area of the support, was unsuccessfully attempted in the first place; the fit was insensible to very large changes in the values assigned to A_S , the coefficients had no statistical significance and, furthermore, it was found that Eq. 3 was overparameterized, as evidenced by dependencies larger than 0.9999. On the other hand, there was a good linear relationship between the variables V_N/V_L and A_L/V_L , as indicated by correlation coefficients higher than 0.997 and by the small coefficients of variation of both the intercepts and the slopes (see second and fifth column in Table 2, respectively). In terms of Eq. 3, since V_L increases almost five times and A_L decreases by about 60% between the lower and the higher stationary phase loading, it is reasonable to conclude that K_S is negligible under the present circumstances. In other words, within the experimental errors of our measurements, the

intercepts and the slopes obtained by means of the linear regression can be identified with the coefficients K_L and K_A , respectively.

These results indicate that hydrocarbon retention in columns containing Fomblin coated on Chromosorb P AW DMCS can be attributed to partition and adsorption on the gas–liquid interface. Depending on the solute, the contribution of adsorption to the total retention ranges between 50–60% at the lower loading and between 10–15% in the column with higher concentration of stationary phase, as calculated by using the K_L and K_A results obtained as indicated in the former paragraph together with the V_L and A_L values characteristic of each column.

Standard enthalpies of solution, corresponding to the transfer of one mole of solute from an ideal vapour phase at a pressure of 1 atm to an hypothetical solution at unitary weight fraction, with behaviour extrapolated from infinite dilution, were calculated by means of the equation [18]

$$\Delta H_L^0 = -R[\partial \ln K_L / \partial (1/T)] - RT(1 - \alpha_2 T) \quad (4)$$

The assumption of ideal vapour behaviour implied in Eq. 4 is justified by a rapid calculation

Table 2

Thermodynamic functions of partition and adsorption of vapours at infinite dilution in Fomblin at 25°C

Solute	$K_L \pm \sigma(K_L)$	$-\Delta H_L^0$ (kJ/mol)	χ^+ (int)	$K_A \pm \sigma(K_A)$ ($\times 10^4$ cm)	$-\Delta H_A^0$ (kJ/mol)	$-\Delta H_C^0$ (kJ/mol)
<i>n</i> -Hexane	29.21 \pm 1.12	24.7 \pm 3.2	2.462	0.573 \pm 0.031	30.5 \pm 5.4	31.6
<i>n</i> -Heptane	65.33 \pm 1.64	28.1 \pm 2.8	2.761	1.239 \pm 0.042	36.4 \pm 4.4	36.6
2-Methylhexane	53.21 \pm 1.31	26.2 \pm 2.8	2.605	0.998 \pm 0.033	35.7 \pm 4.5	34.8
2,3-Dimethylpentane	60.12 \pm 1.71	26.1 \pm 2.7	2.472	0.949 \pm 0.044	35.4 \pm 5.5	34.2
<i>n</i> -Octane	145.1 \pm 6.95	31.1 \pm 4.9	3.058	2.970 \pm 0.177	42.7 \pm 8.2	41.5
2-Methylheptane	117.5 \pm 6.04	30.7 \pm 4.9	2.883	2.337 \pm 0.154	40.4 \pm 8.4	39.7
2,2-Dimethylhexane	88.90 \pm 2.28	28.4 \pm 2.8	2.646	1.681 \pm 0.058	38.7 \pm 4.7	37.3
2,2,4-Trimethylpentane	77.93 \pm 3.10	26.4 \pm 3.7	2.406	1.353 \pm 0.079	38.0 \pm 7.3	35.1
<i>n</i> -Nonane	325.5 \pm 15.7	36.0 \pm 5.6	3.340	6.906 \pm 0.400	45.8 \pm 8.6	46.4
1-Heptene	55.7 \pm 2.13	26.2 \pm 3.8	2.764	1.287 \pm 0.054	36.6 \pm 5.6	36.1
1-Octene	122.9 \pm 3.82	30.6 \pm 3.4	3.057	3.090 \pm 0.097	40.9 \pm 4.6	40.8
Cyclohexane	45.68 \pm 2.03	25.3 \pm 3.4	2.671	0.621 \pm 0.052	30.5 \pm 7.7	33.0
Cyclohexene	48.73 \pm 2.41	26.9 \pm 4.5	2.758	0.872 \pm 0.061	28.8 \pm 6.9	33.3
Methylcyclohexane	79.24 \pm 3.68	27.6 \pm 4.1	2.713	1.165 \pm 0.094	35.1 \pm 8.8	35.4
Ethylcyclohexane	191.1 \pm 9.16	31.0 \pm 4.6	3.020	2.808 \pm 0.233	40.5 \pm 10.2	40.5

[12] which indicates that the non-ideality correction in the case of the most volatile solute at the higher experimental temperature (*n*-hexane at 35°C) amounts to about 0.3% of K_L , a value below our estimation of the experimental error.

Enthalpies of adsorption were computed from the expression

$$\Delta H_A^0 = -R[\partial \ln K_A / \partial (1/T)] \quad (5)$$

and can be shown [19] to correspond to the transfer of one mole of solute from an ideal vapour phase at 1 atm to an ideal adsorbed state in which the molecules of the adsorbate interact with the surface only. In as much as adsorbate–adsorbate interactions are absent, the value of the adsorption enthalpy does not depend on the definition of the adsorbed state; this statement obviously does not apply to the adsorption free energy and entropy.

Solution and adsorption properties have been gathered in Table 2. As already mentioned, the uncertainties assigned to K_L and K_A values are the standard deviations for the intercept and the slope, respectively, obtained in the regression of V_N/V_L against A_L/V_L . The confidence ranges for ΔH_L^0 and ΔH_A^0 at the 95% level were estimated by following a method proposed some years ago [13], which takes into consideration that two successive regressions are necessary to obtain the enthalpy values. The latent heats of condensation, ΔH_C^0 [20], were included in Table 2 for comparison.

The results shown in the table change very regularly with the molecular structure of the solutes. Thus the plots of $\ln K_L$ for *n*-alkanes against their carbon number, *N*, are straight lines with correlation coefficients $r = 0.99997$, and those of $\ln K_A$ against *N* result in $r = 0.9997$. Both K_L and K_A are smaller for branched than for normal alkanes with the same *N*, and (except for the K_L result for 2,3-dimethylpentane) both values decrease as branching increases. The partition coefficients of the alkenes are smaller than those of the corresponding alkanes, while the opposite trend is observed for the adsorption coefficients. Partition coefficients of cycloalkanes are larger than those of the alkanes with the

same carbon number, but the adsorption coefficients of both groups of substances are coincident within experimental error.

As could be expected, less regularity is found when the enthalpies are compared: plots of ΔH_L^0 and of ΔH_A^0 against *N* for *n*-alkanes are well represented by straight lines, but their correlation coefficients drop to 0.994 and 0.991, respectively. In general terms, branched alkanes and alkenes show less negative ΔH_L^0 and ΔH_A^0 values than normal alkanes with the same *N*.

Figs. 4 and 5 show that there is a strong correlation between K_A and K_L on the one hand, and between ΔH_A^0 and ΔH_L^0 on the other. It can be concluded from this behaviour that solution and adsorption are ruled by the same physical factors, either energetic or statistical. However, when the enthalpies of adsorption and of solution are compared with the latent heats of condensation it is found that:

(a) the differences $\Delta H_A^0 - \Delta H_C^0$ are zero within experimental error, indicating that the liquid–gas interface of Fomblin is a low-energy surface on which the hydrocarbon molecules adsorb without experiencing the influence of large repulsion or attraction forces, and

(b) the differences $\Delta H_L^0 - \Delta H_C^0$, rough estimators of infinite dilution excess enthalpies, range between 7 and 10 kJ/mol, indicating that the solute molecules are repelled from the bulk liquid and require the absorption of important quantities of energy in order to penetrate it.

Deviations from the ideal behaviour are more important in bulk mixtures than on surfaces, a fact already detected by Handa and Mukerjee [10], who attributed it to the lower number of nearest neighbours that a molecule has on a surface compared to the situation in bulk solution. These authors semiquantitatively explained the surface properties of *n*-alkane–perfluoroalkane mixtures by assuming unitary surface activity coefficients and modelling the bulk activity coefficients by means of the regular solution and the Flory–Huggins approaches.

Flory–Huggins interaction parameters, χ^* , were calculated with the equation [21]

$$\chi^* = \ln(RT\rho_2/K_L p_1^0 M_1) - \ln(v_1^*/v_2^*) - 1 + (M_1 v_1^*/M_2 V_2^*) \quad (6)$$

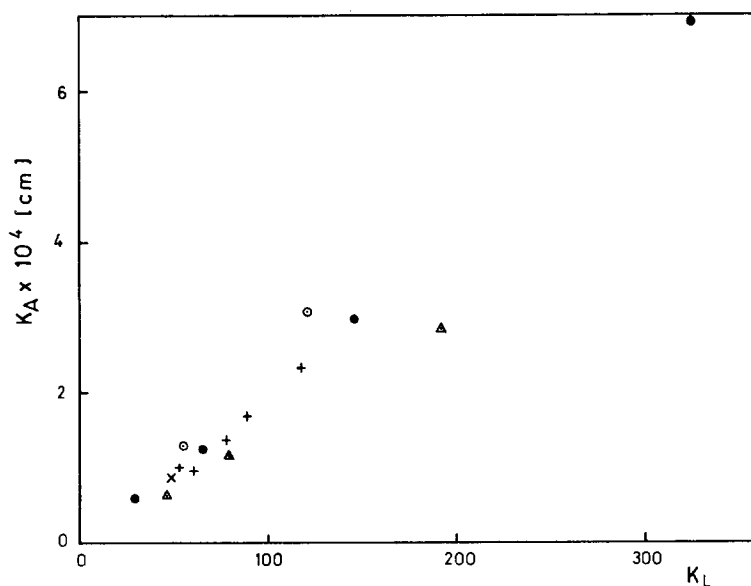


Fig. 4. Adsorption coefficients (K_A) against partition coefficients (K_L) at 25°C. Solutes: (●) *n*-alkanes; (+) branched alkanes; (○) alkenes; (△) cycloalkanes; (×) cyclohexene.

where M_1 and p_1^0 represent the solute molecular mass and vapour pressure at the column temperature, respectively, M_2 is the polymer number averaged molecular mass and v_1^* and v_2^* are

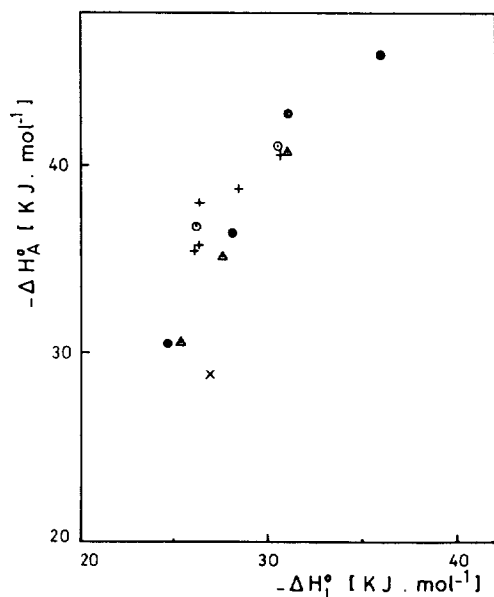


Fig. 5. Enthalpy of adsorption (ΔH_A^0) against enthalpy of solution (ΔH_L^0). Symbols as in Fig. 4.

the solute and the polymer specific “hard core” volumes, calculated by means of the Flory state equation [22] using experimental thermal expansion coefficients. The interaction parameter is a non-dimensional free energy term that includes all the non-combinatorial contributions to excess free energy; Eichinger and Flory [22] express χ^* as a sum of two contributions:

$$\chi^* = \chi^*(fv) + \chi^*(int) \quad (7)$$

where $\chi^*(fv)$ is the free volume contribution, and $\chi^*(int)$ results from contact interactions.

Unusually high χ^* values, ranging between 2.4 and 3.4, are obtained for the systems studied in the present paper. With as only exception *n*-hexane, free volume effects contribute less than 2.5% to the total interaction parameter. The highly positive non-ideality of these mixtures results from very large differences between the force fields surrounding the hydrocarbon and the perfluorocarbon segments.

Values of $\chi^*(int)$ can be found in the fourth column of Table 2; they were fitted by a non-linear regression method to the equation

$$\chi^*(int) = (V_1^*/RT)(\delta_1 - \delta_2)^2 \quad (8)$$

where $V_1^* = M_1 v_1^*$, δ_1 is the solute solubility parameter [23], and δ_2 , the polymer solubility parameter, was the fitting parameter. A value $\delta_2 = 7.65 \text{ J}^{1/2} \text{ cm}^{-3/2}$ is thus obtained; the standard deviation obtained for δ_2 is only $0.084 \text{ J}^{1/2} \text{ cm}^{-3/2}$, indicating that the experimental results fit nicely to Eq. 8. However, the value returned for δ_2 is unusually low and, from our point of view, its physical meaning is not clear. This is by no means a novelty in dealing with polymer solubility parameters.

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